



Geological Survey of Canada Open File 3015 (NTS 63J/13 West and 63K/16)

Exploration Science and Technology Initiative

LAKE SEDIMENT AND WATER GEOCHEMICAL INFILL SURVEY DATA, SNOW LAKE AREA WEST CENTRAL MANITOBA

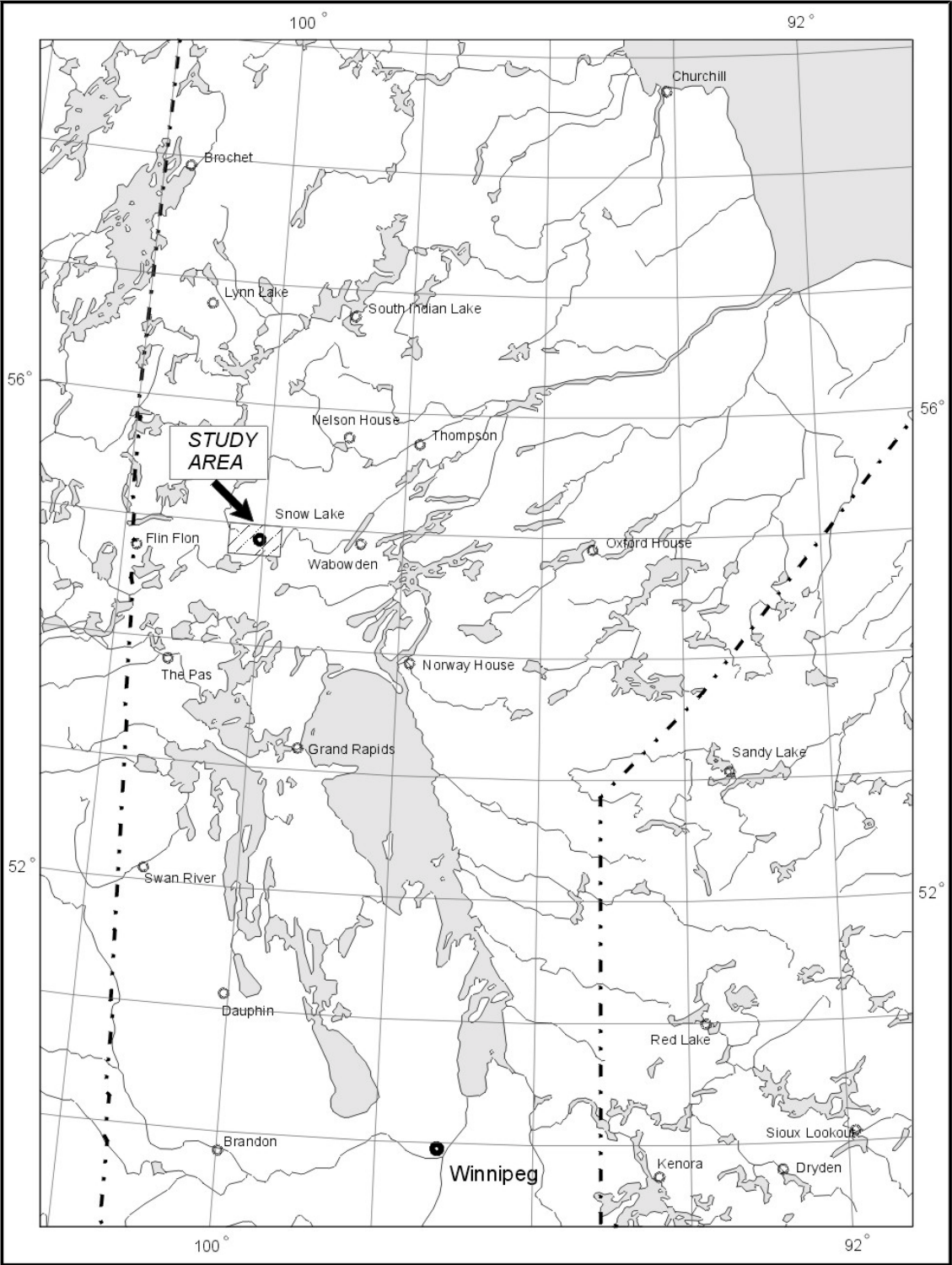


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National Geochemical Reconnaissance

Lake Sediment and Water Geochemical Infill Survey Data

Manitoba 1995



Open File 3015 represents a contribution to the Exploration Science and Technology Initiative (EXTECH).

GSC OPEN FILE 3015
LAKE SEDIMENT AND WATER GEOCHEMICAL DATA
MANITOBA NTS 63J/13 WEST and 63K/16

INTRODUCTION

Open File 3015 presents analytical and statistical data for 35 elements in lake sediments from 346 sample sites in Manitoba. Loss-on-ignition in sediments, and uranium, fluoride, and pH values in waters from these sites are included in this report. Open File 3015 contains geochemical data from an area in the west-central part of the province sampled in 1991 as part of the Exploration Science and Technology Initiative (EXTECH).

The reconnaissance surveys were managed by the Geological Survey of Canada.

Regional geochemical surveys have been carried out by the GSC in western Canada since 1974. A total of 24 open files are available that provide coverage of Precambrian regions of Alberta, Saskatchewan, and Manitoba. Areas surveyed, with associated open file numbers, are shown in Figure 1. Figure 2 illustrates geochemical coverage across Canada. Data from all open files are available on 3.5 or 5.25 inch diskettes, or in the original published form.

Analytical results and field observations are used to build a national geochemical data base for resource assessment, mineral exploration, geological mapping and environmental studies. Sample collection, preparation procedures and analytical methods are strictly specified and carefully monitored to ensure consistent and reliable results regardless of the area, the year or the analytical laboratory.

CREDITS

P.W.B. Friske coordinated the activities of GSC staff.

Contracts were let to the following companies for preparation and analysis, and were managed by Geological Survey of Canada staff as follows:

Preparation: Bondar-Clegg & Company Limited
Gloucester, Ontario
J.J. Lynch (GSC)

Analysis: Becquerel Laboratories
Mississauga, Ontario

Cantech Laboratories Inc.
Calgary, Alberta

Bondar-Clegg & Company Limited
Gloucester, Ontario

J.J. Lynch (GSC)

M. McCurdy edited open files and coordinated open file production.

H. Gross provided computer processing support.

C.C. Durham and S.J. Day provided technical assistance.

DESCRIPTION OF SURVEY AND SAMPLE MANAGEMENT

Helicopter-supported sample collection was carried out during the fall of 1991. Lake sediment and water samples were collected at an average density of one sample per 3.87 sq. km throughout the 1 340 square kilometres covered by the survey.

Samples were arranged in groups (blocks) of twenty. Each group of twenty contained site duplicate samples, that is, two samples from a single site; the group also contained an analytical duplicate sample pair (a single site sample split and placed in two non-adjacent sample vials). Finally, each group included a control reference sample. The functions of these samples are described in the section titled, Presentation and Interpretation of Gold Data.

Field observations were recorded on standard forms used by the Geological Survey of Canada (Garrett, 1974).

Site positions were marked on 1:50 000 scale NTS maps in the field and later digitized at the Geological Survey of Canada in Ottawa to obtain Universal Transverse Mercator (UTM) coordinates. The dominant rock types in the lake catchment basins were identified on appropriate geological maps used as the bedrock geological base on NGR maps.

In Ottawa, field dried samples were air-dried and crushed: particle reduction was accomplished using a ceramic puck mill. The minus 80 mesh (177 micron) fraction was obtained and used for subsequent analyses. At this time, control reference and blind (analytical) duplicate samples were inserted into each block of twenty sediment samples. For the water samples, only control reference samples were inserted into the block. There were no blind duplicate water samples.

Analytical data from labs were monitored for reliability with standard methods used by the Applied Geochemistry Subdivision at the GSC.

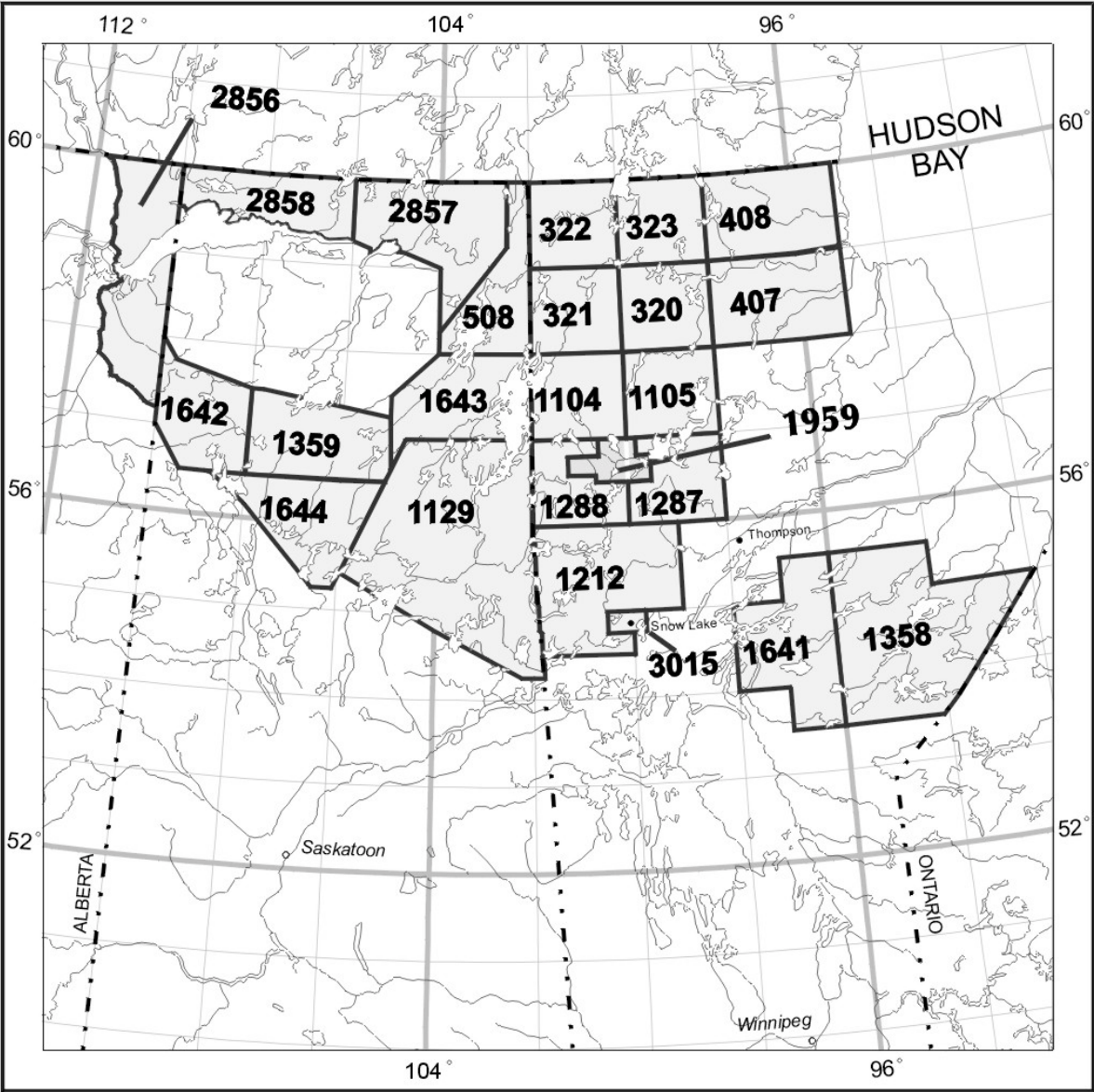


Fig. 1. Lake sediment survey coverage of Alberta, Saskatchewan, and Manitoba, showing current GSC open file numbers.

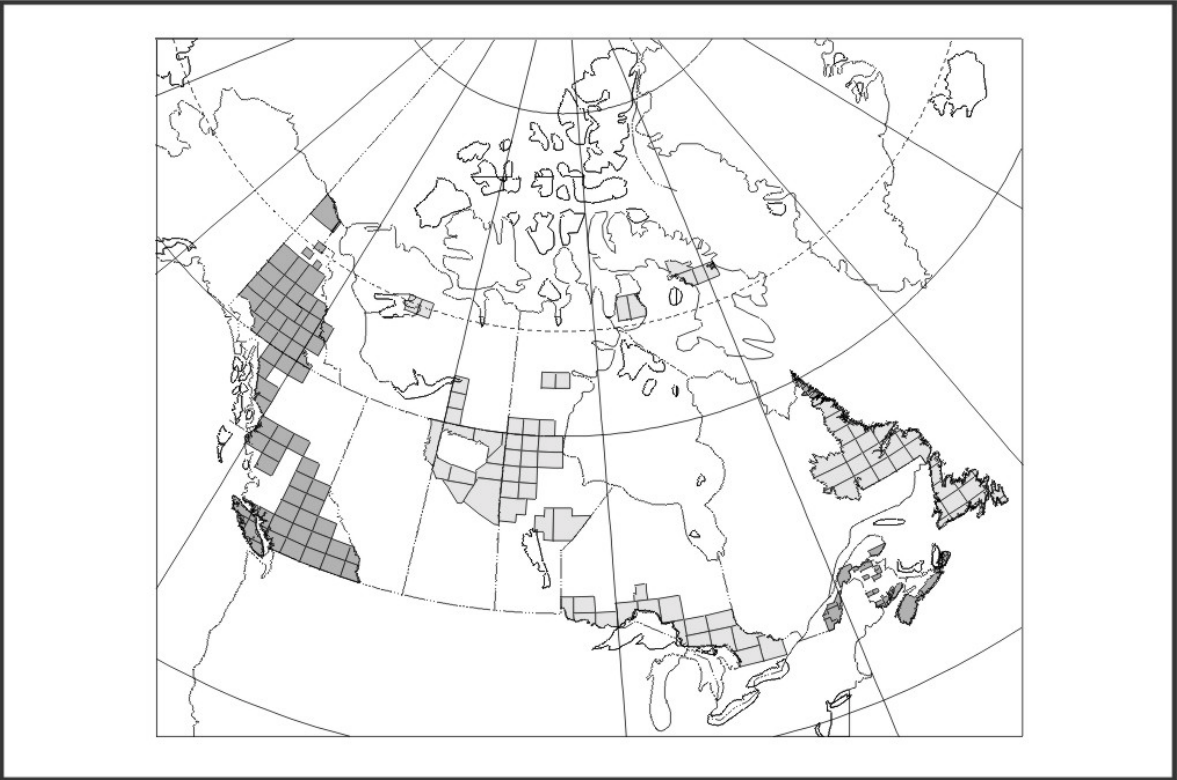


Fig. 2. Drainage surveys to National Geochemical Reconnaissance standards.

ANALYTICAL PROCEDURES

Instrumental Neutron Activation Analysis (INAA)

Weighed and encapsulated samples are packaged for irradiation along with internal standards and international reference materials. Samples and standards are irradiated together with neutron flux monitors in a two-megawatt pool-type reactor. After a seven day decay period, samples are measured on a high resolution germanium detector. Computer control is achieved with a Microvax II computer. Typical counting times are 500 seconds. Elements determined by INAA include: Ag, As, Au, Ba, Br, Cd, Ce, Co, Cr, Cs, Eu, Fe, Hf, Ir, La, Lu, Mo, Na, Ni, Rb, Sb, Sc, Se, Sm, Sn, Ta, Tb, Te, Th, U, W, Yb, Zn, and Zr. The sample weights are also reported. Data for Ag, Cd, Ir, Mo, Ni, Se, Sn, Te, Zn, and Zr are not published because of inadequate detection limits and/or precision.

Atomic Absorption Spectroscopy (AAS) and Other Analyses

For the determination of Zn, Cu, Pb, Ni, Co, Ag, Mn, Fe, and Cd, a 1 g sample is reacted with 6 mL of a mixture of 4M HNO₃ and M HCl in a test tube over night at room temperature. After digestion, the test tube is immersed in a hot water bath at room temperature and brought up to 90 degrees C and held at this temperature for 2 hours with periodic shaking. The sample solution is then diluted to 20 mL with metal-free water and mixed. Zn, Cu, Pb, Ni, Co, Ag, Mn, Fe and Cd are determined by atomic absorption spectroscopy using an air-acetylene flame. Background corrections are made for Pb, Ni, Co, Ag and Cd.

Molybdenum and vanadium are determined by atomic absorption spectroscopy using a nitrous oxide acetylene flame. A 0.5 g sample is reacted with 1.5 mL concentrated HNO₃ in a test tube overnight at room temperature. After digestion, the test tube is immersed in a hot water bath at room temperature and brought up to 90 degrees C and held at this temperature for 30 minutes with periodic shaking. At this point, 0.5 ml concentrated HCl is added and the digestion continued at 90 degrees C for an additional 90 minutes. After cooling, 8 ml of 1250 ppm Al solution are added and the sample solution diluted to 10 ml before aspiration.

Mercury is determined by the Hatch and Ott procedure with some modifications. The method is described by Jonasson et al. (1973). A 0.5 g sample is reacted with 20 mL concentrated HNO₃ and 1 mL concentrated HCl in a test tube for 10 minutes at room temperature prior to 2 hours of digestion with mixing at 90 degrees C in a hot water bath. After digestion, the sample solutions are cooled and diluted to 100 mL with metal-free water. The Hg present is reduced to the elemental state by the addition of 10 mL 10% w/v SnSO₄ in M H₂SO₄. The Hg vapour is then flushed by a

stream of air into an absorption cell mounted in the light path of an atomic absorption spectrophotometer. Absorption measurements are made at 253.7 nm.

Loss-on-ignition is determined using a 500 mg sample. The sample, weighed into 30 mL beaker, is placed in a cold muffle furnace and brought up to 500 degrees C over a period of 2 -3 hours. The sample remains at this temperature for 4 hours and then allowed to cool to room temperature for weighing.

Fluorine is determined in lake sediments as described by Ficklin (1970). A 250 mg sample is sintered with 1 g of a flux consisting of two parts by weight sodium carbonate and one part by weight potassium nitrate. The residue is then leached with water. The sodium carbonate is neutralized with 10 mL 10% (w/v) citric acid and the resulting solution is diluted to 100 mL with water. The fluoride in the resulting solution is then measured using a fluoride ion electrode. Standard solutions contain sodium carbonate and citric acid in the same quantities as the sample solution.

Water Analyses

Fluoride in lake water samples is determined using a fluoride electrode. Prior to measurement, an aliquot of the sample is mixed with an equal volume of TISAB II buffer solution (total ionic strength adjustment buffer). The TISAB II buffer solution is prepared as follows: to 50 mL metal-free water add 57 mL glacial acetic acid, 58 g NaCl and 4 g CDTA (cyclohexylene dinitrilo tetraacetic acid). Stir to dissolve and cool to room temperature. Using a pH meter, adjust the pH between 5.0 and 5.5 by slowly adding 5 M NaOH solution. Cool and dilute to one litre in a volumetric flask.

Hydrogen ion activity (pH) is measured with a combination glass-calomel electrode and a pH meter.

Uranium in waters is determined by a laser-induced fluorometric method using a Scintrex UA-3 uranium analyzer. A complexing agent, known commercially as Fluran and composed of sodium pyrophosphate and sodium monophosphate (Hall, 1979) is added to produce the uranyl pyrophosphate species which fluoresces when exposed to the laser. Since organic matter in the sample can cause unpredictable behaviour, a standard addition method is used. Further, the reaction of uranium with Fluran can be delayed or sluggish; for this reason an arbitrary 24 hour time delay between the addition of the Fluran and the actual reading is incorporated into this method. In practice, 500 microlitres of Fluran solution are added to a 5 mL sample and allowed to stand for 24 hours. At the end of this period fluorescence readings are made with the addition of 0.0, 0.2 and 0.4 ppb U. For high samples the additions are 0.0, 2.0 and 4.0 (20

microlitre aliquots of either 55 or 550 ppb U are used). All readings are taken against a sample blank.

COMPARISON OF DATA PRODUCED BY TWO METHODS

The data listed allows users to make a comparison of data generated by two different analytical methods for a couple of elements. Before attempting such a comparison some caution should be exercised.

The 'wet chemistry' data for Co and Fe were obtained by AAS using a partial extraction (HNO₃ and HCl). The data for these elements obtained by INAA produces 'total' data. Hence, the 'wet chemistry' data will likely be somewhat lower than the INAA data.

PRESENTATION AND INTERPRETATION OF GOLD DATA

The following general discussion reviews the format used to present the gold geochemical data and outlines some important points to consider when interpreting this data. This discussion is included in recognition of the special geochemical behaviour and mode of occurrence of gold in nature and the resultant difficulties in obtaining and analyzing samples which reflect the actual concentration level at a given site.

Samples that have gold values that are statistically above approximately the 90th percentile, or those with LOI values below 10%, are normally analyzed again in accordance with standard NGR procedures. **There will be no repeat data published in Open File 3015 however, as insufficient material remained after the initial neutron activation analyses.** The correct interpretation of gold geochemical data from regional stream sediment or lake sediment surveys requires an appreciation of the unique chemical and physical characteristics of gold and its mobility in the surficial environment. Key properties of gold that distinguish its geochemical behaviour from most other elements (Harris, 1982) include:

- 1) Gold occurs most commonly in the native form which is chemically and physically resistant. A significant proportion of the metal is dispersed in a micron-sized particulate form, and the high specific gravity of gold results in a heterogeneous distribution, especially in stream sediment and clastic-rich (low LOI) lake sediment environments. In organic-rich fluvial and lake sediments, gold distribution appears to be more homogeneous.

- 2) Gold typically occurs at low concentrations in the ppb range. Whereas gold concentrations of only a few ppm may represent economic deposits, background levels in stream and centre-lake sediments seldom exceed 10 ppb, and commonly are near the detection limit of 2 ppb.

These factors result in a particle sparsity effect wherein very low concentrations of gold are heterogeneously enriched or depleted in the surficial environment. Hence, a major problem facing the geochemist is to obtain a representative sample. In general, in areas where concentrations of gold in sediments are low, and/or grain sizes of the gold present relatively high, proportionally larger samples are required to reduce the uncertainty between subsample analytical values and actual values. Conversely, as actual gold concentrations increase or grain size decreases, the number of gold particles to be shared in random subsamples increases and variability of results decreases (Clifton et al., 1969; Harris, 1982). The limited amount of material collected during the rapid, reconnaissance-style regional surveys and the need to analyze for a broad spectrum of elements, precludes the use of a significantly large sample weight for the gold analyses. Therefore, to obtain representative samples, grain size is reduced by sieving and milling of the dried sediments. The following control methods are currently employed to evaluate and monitor the sampling and analytical variability which are inherent in the determination of gold as well as other elements in geochemical media:

- (1) For each block of 20 samples:
 - (a) random insertion of a standard reference sample to control analytical accuracy and long-term precision;
 - (b) collection of a field duplicate (two samples from one site) to measure sampling and analytical variance;
 - (c) analysis of a second subsample (blind duplicate) from one sample to measure and control short-term precision or analytical variance.

In summary, geochemical follow-up investigations for gold should be based on a careful consideration of all geological and geochemical information, and especially a careful appraisal of gold geochemical data and its variability. In some instances, prospective follow-up areas may be indirectly identified by pathfinder element associations in favourable geology, even if an analogous gold response due to natural variability is not present.

Table 1. Summary of Analytical Data and Methods.

ELEMENT	DETECTION LEVEL	METHOD
<u>SEDIMENTS:</u>		
Ag Silver	0.2 ppm	AAS
As Arsenic	0.5 ppm	INAA
Au Gold	2 ppb	INAA
AuWt Sample Weight	0.01 g	-
Ba Barium	50 ppm	INAA
Br Bromine	0.5	INAA
Cd Cadmium	0.2	AAS
Ce Cerium	5	INAA
Co Cobalt	2	AAS
Co Cobalt	5	INAA
Cr Chromium	20	INAA
Cs Cesium	0.5	INAA
Cu Copper	2	AAS
Eu Europium	1	InAA
F Fluorine	40	ISE
Fe Iron	0.02	AAS
Fe Iron	0.2	INAA
Hf Hafnium	1	INAA
Hg Mercury	10	CV-AAS
La Lanthanum	2	INAA
LOI Loss-on-ignition	1.0	GRAV
Lu Lutecium	0.2	INAA
Mn Manganese	5	AAS
Mo Molybdenum	2	AAS
Na Sodium	0.02	INAA
Ni Nickel	2	AAS
Pb Lead	2	AAS
Rb Rubidium	5	INAA
Sb Antimony	0.1	INAA
Sc Scandium	0.2	INAA
Sm Samarium	0.1	INAA
Ta Tantalum	0.5	INAA
Tb Terbium	0.5	INAA
Th Thorium	0.2	INAA
U Uranium	0.2	INAA
V Vanadium	5	AAS
W Tungsten	1	IMAA
Yb Ytterbium	1	IMAA
Zn Zinc	2	AAS
<u>WATERS:</u>		
F-W Fluoride	20 ppb	ISE
pH Hydrogen ion activity	- -	GCM
U-W Uranium	0.05 ppb	LIF

- AAS - atomic absorption spectrometry
CV-AAS - cold vapour / atomic absorption spectrometry
GCM - glass Calomel electrode and pH meter
GRAV - gravimetry
INAA - Instrumental Neutron Activation Analysis
ISE - ion selective electrode
LIF - laser-induced fluorescence

Once an anomalous area has been identified, field investigations should be designed to include detailed geochemical follow-up surveys and collection of large representative samples. Subsequent repeat subsample analyses will increase the reliability of results and permit a better understanding of natural variability which can then be used to improve sampling methods and interpretation.

Section of the Geological Association of Canada, pp. 53-67.

Jonasson, I.R., Lynch, J.J. and Trip, L.J. (1973) Field and laboratory methods used by the Geological Survey of Canada in geochemical surveys; No. 12, Mercury in Ores, Rocks, Soils, Sediments and Water, Geol. Surv. Can. Paper 73-21.

FIELD DATA LEGEND

Table 2 describes the field and map information appearing on the following pages preceding the analytical data for each sample site.

REFERENCES

Clifton, H.E., Hunter, R.E., Swanson, F.J., and Phillips, R.L. (1969). Sample size and meaningful gold analysis; U.S. Geological Survey Professional Paper 625-C.

Ficklin, W.H. (1970). A rapid method for the determination of fluoride in rocks and soils, using an ion selective electrode; U.S. Geol. Surv. Paper 700C, pp. 186-188.

Friske, P.W.B. and Hornbrook, E.H.W. (1991) Canada's National Geochemical Reconnaissance Program; in Transactions of the Institution of Mining and Metallurgy, Section B; Volume 100, p. 47-56.

Fedikow, M.A., Athayde, P., and Galley, A.G. (1989) MDS Map 14, Mineral deposits and occurrences in the File Lake area, Manitoba (NTS 63J/13). To accompany Report No. 14, Mineral Deposits Series, Manitoba Department of Mines and Energy; 1:50 000.

Fedikow, M.A., Ostry, G., Ferreira, K.J, and Galley, A.G. (1988) Mineral deposits and occurrences in the File Lake area, NTS 63K/16: Manitoba Energy and Mines, Mineral Deposit Series Report No. 5; 1:50 000.

Garrett, R.G. (1974) Field data acquisition methods for applied geochemical surveys at the Geological Survey of Canada; Geol. Surv. Can. Paper 74-52.

Hall, G.E.M. (1979) A study of the stability of uranium in waters collected from various geological environments in Canada; in Current Research, Part A, Geological Survey of Canada Paper 79-1A, pp. 361-365.

Harris, J.F. (1982) Sampling and analytical requirements for effective use of geochemistry in exploration for gold; in Levinson, A.A., Editor, Precious Metals in the Northern Cordillera, proceedings of a symposium sponsored by the Association of Exploration Geochemists and the Cordilleran

TABLE 2. Field Observations Legend

FIELD RECORD	DEFINITION	TEXT CODE
MAP	National Topographic System (NTS): lettered quadrangle (1:250 000 or 1:50 000 scale) Part of sample number	63J/13 or 63K/16
SAMPLE ID	Remainder of sample number: Year..... Field crew..... Sample sequence number.....	91 1 or 3 001-999
REP STAT	Replicate status; the relationship of the sample to others within the analytical block of 20: Routine regional sample First of field duplicate Second of field duplicate	
ZN	Zone (7 to 22)	
UTM	Universal Transverse Mercator (UTM) Coordinate System; digitized sample location coordinates.	
EASTING	UTM Easting in metres.	
NORTHING	UTM Northing in metres.	
ROCK UNIT	Major rock type of catchment area: INTRUSIONS pegmatites..... felsic intrusions; granite, granodiorite, tonalite. mafic intrusives; gabbro, diorite, ultramafics..... KISSEYNEW GNEISS TERRAIN quartzofeldspathic gneiss and migmatite..... gneissic metagreywacke and migmatite..... amphibolite and hornblende-biotite-quartz-plagioclase gneiss..... MISSI GROUP felsic volcanics..... mafic volcanics..... sandstone and conglomerate..... AMISK GROUP greywacke, siltstone, mudstone..... felsic volcanic rocks..... mafic to intermediate volcanic rocks..... unclassified.....	13 12 11 10 9 8 7 6 5 4 3 2 1
ROCK AGE	Predominant age of rock type in catchment area: Proterozoic.....	04
LAKE AREA	The area of the water body sampled: Pond..... ¼ to 1 square kilometre..... 1 to 5 square kilometres..... greater than 5 square kilometres.....	Pond .25-1 1-5 >5
LAKE DEPTH	Distance in metres from the surface of the lake to the bottom.....	0 - 99
TERRAIN RELIEF	Relief of lake catchment basin: Low..... Medium..... High.....	Low Med Hi

FIELD RECORD	DEFINITION	TEXT CODE
SAMPLE CONT	Contamination; human or natural: None..... Work..... Camp..... Fuel..... Gossan.....	- Wo Ca Fu Go
SAMPLE COLOUR	Sediment sample colour; up to two colours may be selected: Tan..... Yellow..... Green..... Grey..... Brown..... Black.....	Tan Yellow Green Grey Brown Black
SUSPEND MATL	Suspended matter in water: None..... Light..... Heavy.....	- Light Heavy
Miscellaneous	Refers to mising data in any field no sample material for analysis parts per million parts per billion percent weight (of sample) gram	* ns ppm ppb pct Wt g